

Mole & Avogadro's Number Worksheet

Complete the following table with the appropriate information.

$$[\text{Avogadro's Number } (N_A) = 6.023 \times 10^{23}]$$

Weight	Moles	Molecules
25 g of NaCl		
146 g of NaCl		
49.05 g of H ₂ SO ₄		
125 g of H ₂ SO ₄		
100 g of KMnO ₄		
269 g of KMnO ₄		
74 g of KCl		
35 g of CuSO ₄ .5H ₂ O		
18.7 g of KCl		
799 g of CuSO ₄ .5H ₂ O		

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Answers

[Avogadro's Number (N_A) = 6.023×10^{23}]

Weight	Moles	Molecules
25 g of NaCl	0.43	2.58×10^{23}
146 g of NaCl	2.5	1.5×10^{24}
49.05 g of H_2SO_4	0.5	3.011×10^{23}
125 g of H_2SO_4	1.27	7.64×10^{23}
100 g of $KMnO_4$	0.633	3.81×10^{23}
269 g of $KMnO_4$	1.7	1.02×10^{24}
74 g of KCl	0.99	5.96×10^{23}
35 g of $CuSO_4 \cdot 5H_2O$	0.14	8.4×10^{22}
18.7 g of KCl	0.25	1.5×10^{23}
799 g of $CuSO_4 \cdot 5H_2O$	3.2	1.9×10^{24}